

	Henry's law constant equilibrium reaction	Temp. dep. at 298 K (moles L ⁻¹ atm ⁻¹)	$-\frac{dH_h}{R}$ (K)
H_{HNO_3}	$\text{HNO}_3(\text{g}) \leftrightarrow \text{HNO}_3(\text{aq})$	2.10×10^5	–
$H_{\text{H}_2\text{O}_2}$	$\text{H}_2\text{O}_2(\text{g}) \leftrightarrow \text{H}_2\text{O}_2(\text{aq})$	7.45×10^4	7300
H_{NH_3}	$\text{NH}_3(\text{g}) \leftrightarrow \text{NH}_3 * \text{H}_2\text{O}$	62	4110
H_{SO_2}	$\text{SO}_2(\text{g}) \leftrightarrow \text{SO}_2 * \text{H}_2\text{O}$	1.23	3150
H_{CO_2}	$\text{CO}_2(\text{g}) \leftrightarrow \text{CO}_2 * \text{H}_2\text{O}$	3.40×10^{-2}	2440
H_{O_3}	$\text{O}_3(\text{g}) \leftrightarrow \text{O}_3(\text{aq})$	1.13×10^{-2}	2540