

	Equilibrium reaction	Dissociation constant at 298K (moles L ⁻¹)	Temp. dep. $\frac{-dH_d}{R}$ (K)
K_{HNO_3}	$\text{HNO}_3(\text{aq}) \leftrightarrow \text{H}^+ + \text{NO}_3^-$	15.4	8700
K_{SO_2}	$\text{SO}_2 * \text{H}_2\text{O} \leftrightarrow \text{H}^+ + \text{HSO}_3^-$	1.3×10^{-2}	1960
K_{NH_3}	$\text{NH}_3 * \text{H}_2\text{O} \leftrightarrow \text{NH}_4^+ + \text{OH}^-$	1.7×10^{-5}	-450
K_{CO_2}	$\text{CO}_2 * \text{H}_2\text{O} \leftrightarrow \text{H}^+ + \text{HCO}_3^-$	4.3×10^{-7}	-1000
K_{HSO_3}	$\text{HSO}_3^- \leftrightarrow \text{H}^+ + \text{SO}_3^{2-}$	6.6×10^{-8}	1500
K_{HCO_3}	$\text{HCO}_3^- \leftrightarrow \text{H}^+ + \text{CO}_3^{2-}$	4.68×10^{-11}	-1760
	$\text{H}_2\text{SO}_4(\text{aq}) \leftrightarrow \text{H}^+ + \text{HSO}_4^-$	∞	-
$K_{\text{H}_2\text{SO}_4}$	$\text{HSO}_4^- \leftrightarrow \text{H}^+ + \text{SO}_4^{2-}$	1.2×10^{-2}	2720